

SOLUBILITY CURVES

Name _____

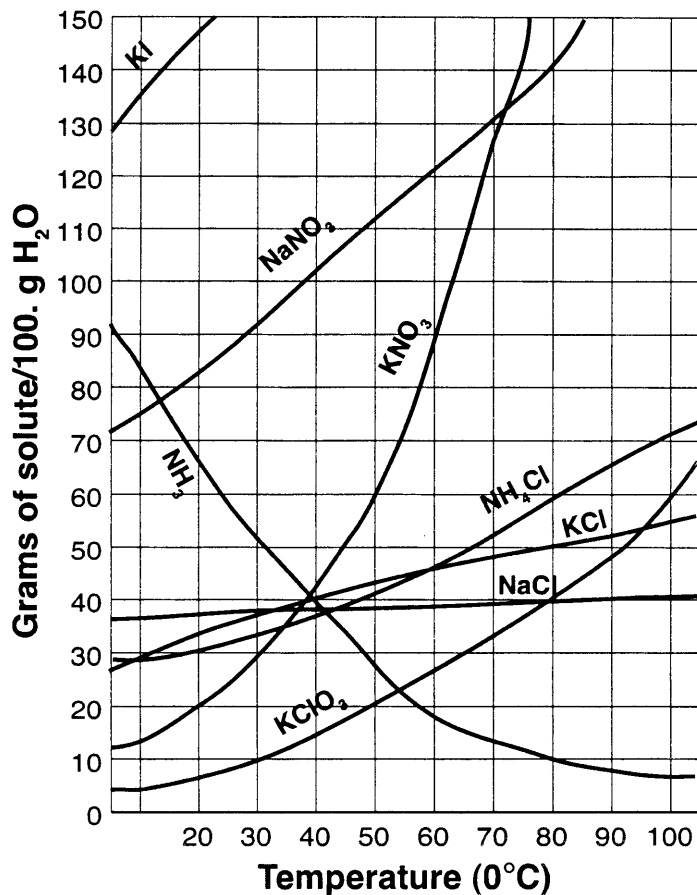
Answer the following questions based on the solubility curve below.

- Which salt is least soluble in water at 20° C? _____
- How many grams of potassium chloride can be dissolved in 200 g of water at 80° C?

- At 40° C, how much potassium nitrate can be dissolved in 300 g of water? _____
- Which salt shows the least change in solubility from 0° – 100° C?

- At 30° C, 90 g of sodium nitrate is dissolved in 100 g of water. Is this solution saturated, unsaturated or supersaturated?

- A saturated solution of potassium chlorate is formed from one hundred grams of water. If the saturated solution is cooled from 80° C to 50° C, how many grams of precipitate are formed? _____
- What compound shows a decrease in solubility from 0° to 100° C? _____
- Which salt is most soluble at 10° C? _____
- Which salt is least soluble at 50° C? _____
- Which salt is least soluble at 90° C? _____



MOLARITY (M)

Name _____

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}}$$

Solve the problems below.

1. What is the molarity of a solution in which 58 g of NaCl are dissolved in 1.0 L of solution?

2. What is the molarity of a solution in which 10.0 g of AgNO₃ is dissolved in 500. mL of solution?

3. How many grams of KNO₃ should be used to prepare 2.00 L of a 0.500 M solution?

4. To what volume should 5.0 g of KCl be diluted in order to prepare a 0.25 M solution?

5. How many grams of CuSO₄•5H₂O are needed to prepare 100. mL of a 0.10 M solution?

MOLARITY BY DILUTION

Name _____

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid = $M \times V$ then, $M_1 \times V_1 = M_2 \times V_2$. Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

4. To how much water should 50. mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

5. To how much water should 100. mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

MOLALITY (m)

Name _____

$$\text{Molality} = \frac{\text{moles of solute}}{\text{Kg of solvent}}$$

Solve the problems below.

1. What is the molality of a solution in which 3.0 moles of NaCl is dissolved in 1.5 Kg of water?

2. What is the molality of a solution in which 25 g of NaCl is dissolved in 2.0 Kg of water?

3. What is the molality of a solution in which 15 g of I₂ is dissolved in 500. g of alcohol?

4. How many grams of I₂ should be added to 750 g of CCl₄ to prepare a 0.020 m solution?

5. How much water should be added to 5.00 g of KCl to prepare a 0.500 m solution?

NORMALITY (N)

Name _____

normality = molarity x total positive oxidation number of solute

Example: What is the normality of 3.0 M of H_2SO_4 ?

Solution: Since the total positive oxidation number of H_2SO_4 is +2 (2 H^+), $N = 6.0$.

Solve the problems below.

1. What is the normality of a 2.0 M NaOH solution?

2. What is the normality of a 2.0 M H_3PO_4 solution?

3. A solution of H_2SO_4 is 3.0 N. What is its molarity?

4. What is the normality of a solution in which 2.0 g of $\text{Ca}(\text{OH})_2$ is dissolved in 1.0 L of solution?

5. How much AlCl_3 should be dissolved in 2.00 L of solution to produce a 0.150 N solution?

ELECTROLYTES

Name _____

Electrolytes are substances that break up (dissociate or ionize) in water to produce ions. These ions are capable of conducting an electric current.

Generally, electrolytes consist of acids, bases and salts (ionic compounds). Nonelectrolytes are usually covalent compounds, with the exception of acids.

Classify the following compounds as either an electrolyte or a nonelectrolyte.

Compound	Electrolyte	Nonelectrolyte
1. NaCl		
2. CH ₃ OH (methyl alcohol)		
3. C ₃ H ₅ (OH) ₃ (glycerol)		
4. HCl		
5. C ₆ H ₁₂ O ₆ (sugar)		
6. NaOH		
7. C ₂ H ₅ OH (ethyl alcohol)		
8. CH ₃ COOH (acetic acid)		
9. NH ₄ OH (NH ₃ + H ₂ O)		
10. H ₂ SO ₄		

EFFECT OF A SOLUTE ON FREEZING AND BOILING POINTS

Name _____

We use the following formulas to calculate changes in freezing and boiling point due to the presence of a nonvolatile solute. Freezing point is always lowered, boiling point is always raised.

$$\Delta T_F = m \times \text{d.f.} \times k_F$$

$$k_B \text{H}_2\text{O} = 0.52^\circ \text{C/m}$$

$$\Delta T_B = m \times \text{d.f.} \times k_B$$

$$k_F \text{H}_2\text{O} = 1.86^\circ \text{C/m}$$

m = molality of solution

k_F and k_B = constants for particular solvent

d.f. = dissociation factor (how many particles solute breaks up into:
for a nonelectrolyte, d.f. = 1)

(Theoretical Dissociation Factor is always greater than observed effect.)

Solve the problems below.

1. What is the new boiling point if 25 g of NaCl is dissolved in 1.0 Kg of water?

2. What is the freezing point of the solution in Problem 1?

3. What are the new freezing and boiling points of water if 50. g of ethylene glycol (molecular mass = 62 g/mol) is added to 50. g of water?

4. When 5.0 g of a nonelectrolyte is added to 25 g of water, the new freezing point is -2.5°C . What is the molecular mass of the unknown compound?

SOLUBILITY (POLAR VS. NONPOLAR)

Name _____

Generally, "like dissolves like." Polar molecules dissolve other polar molecules and ionic compounds. Nonpolar molecules dissolve other nonpolar molecules. Alcohols, which have characteristics of both, tend to dissolve in both types of solvents, but will not dissolve ionic solids.

Check the appropriate columns as to whether the solute is soluble in a polar or nonpolar solvent.

SOLUTES	SOLVENTS		
	Water	CCl ₄	Alcohol
1. NaCl			
2. I ₂			
3. ethanol			
4. benzene			
5. Br ₂			
6. KNO ₃			
7. toluene			
8. Ca(OH) ₂			