

Unit 1 – Atomic Structure

4.1 Defining the Atom

I. Atomic Theory

A. Modern Atomic Theory

1. All matter is made up of very tiny particles called atoms
2. Atoms of the same element are chemically alike
3. Individual atoms of an element may not all have the same mass. However, the atoms of an element have a definite average mass that is characteristic of the element
4. Atoms of different elements have different average masses
5. Atoms are not subdivided, created, or destroyed in chemical rxns

II. Sizes of Atoms

A. Atomic radius

1. 40 to 270 picometers (pm)
 - a. $1 \text{ pm} = 10^{-12} \text{ m}$
2. Most of the atomic radius is due to the electron cloud

B. Nuclear radius

1. 0.001 pm
2. density is 2×10^8 metric tons/cm³
 - a. 1 metric ton = 1000kg

4.2 Structure of the Nuclear Atom

I. The Electron

A. Discovery

1. Joseph John Thomson (1897)
 - a. Cathode ray tube produces a ray with a constant charge to mass ratio
 - b. All cathode rays are composed of identical negatively charged particles (electrons)

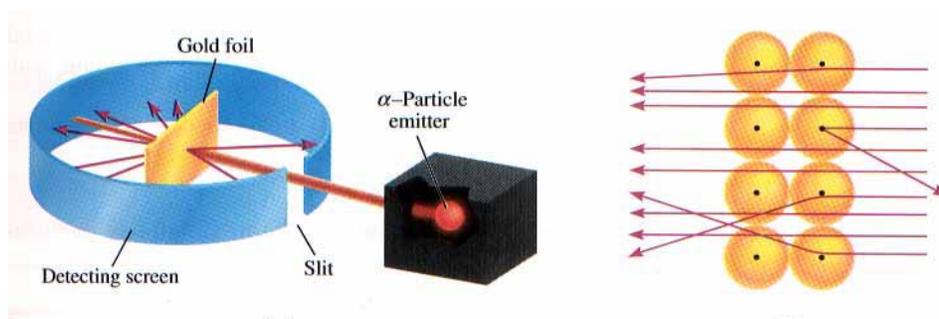
B. Inferences from the properties of electrons

1. Atoms are neutral, so there must be positive charges to balance the negatives
2. Electrons have little mass, so atoms must contain other particles that account for most of the mass

II. The Nucleus

A. The Rutherford Experiment (1911)

1. Alpha particles (helium nuclei) fired at a thin sheet of gold
 - a. Assumed that the positively charged particles were bounced back if they approached a positively charged atomic nucleus head-on (Like charges repel one another)



2. Very few particles were greatly deflected back from the gold sheet
 - a. nucleus is very small, dense and positively charged
 - b. most of the atom is empty space

III. Structure of the Nucleus

1. Protons
 - a. Positive charge, mass of 1.673×10^{-27} kg
 - b. The number of protons in the nucleus determines the atom's identity and is called the atomic number
2. Neutrons
 - a. No charge, mass of 1.675×10^{-27} kg
3. Nuclear Forces
 - a. Short range attractive forces:
neutron-to-neutron, proton-to-proton, proton-to-neutron

| Particle | Symbols | Relative charge | Mass Number |
|----------|--------------------|-----------------|-------------|
| Electron | e^- ${}_{-1}^0e$ | -1 | 0 |
| Proton | p^+ ${}_{1}^1H$ | +1 | 1 |
| Neutron | n^0 ${}_{0}^1n$ | 0 | 1 |

4.3 Distinguishing Among Atoms

I. Atomic Number, Mass Number, and Isotopes

A. Atomic Number (Z)

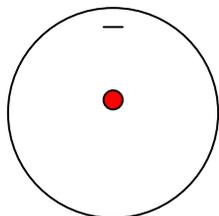
1. The number of protons in the nucleus of each atom of that element
2. Atoms are identified by their atomic number
3. Because atoms are neutral,
protons = # electrons
4. Periodic Table is in order of increasing atomic number

B. Mass Number

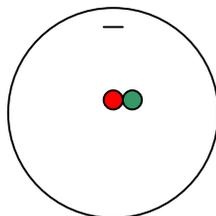
1. The total number of protons and neutrons in the nucleus of an isotope

C. Isotopes

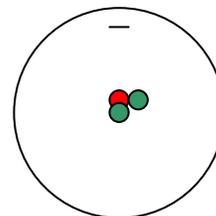
1. Atoms of the same element that have different masses
2. All elements of the same element have the same # of protons, but may vary in the number of neutrons
3. Although isotopes have different masses, they do not differ significantly in their chemical behavior
4. Hydrogen as an example:



Protium (hydrogen-1)



Deuterium (hydrogen-2)



Tritium (hydrogen-3)

D. Designating Isotopes

1. Hyphen notation
 - a. Mass number is written after the name of the element
(1) hydrogen-2, helium-4
2. Nuclear Symbol
 - a. Composition of the nucleus using the element's symbol
(1) ${}_{1}^2H$ Mass number = 2 Atomic number = 1
(2) ${}_{2}^4He$ Mass number = 4 Atomic number = 2

II. Using Atomic Mass

A. Average Atomic Masses

1. The weighted average of the atomic masses of the naturally occurring isotopes of an element
 - a. Atomic masses on the periodic table are average masses
 - b. In calculations using atomic mass, we will round the masses to two decimal places before doing calculations

Examples: Mg = 24.3050 we use: 24.31
 O = 15.9994 we use: 16.00
 N = 14.00674 we use: 14.01

III. Avogadro's Number and the Mole

A. The Mole

1. The amount of substance that contains as many particles as there are in exactly 12 grams of carbon-12
2. The amount of substance that contains the Avogadro number of particles

B. Avogadro's Number

1. The number of particles in exactly one mole of a pure substance
2. Avogadro's number = 6.022×10^{23}

C. Molar Mass

1. The mass of one mole of a pure substance
 - a. Units are grams/mole (or g/mol)
 - b. Molar mass of an element equals the average atomic mass in gram units

25.1 Nuclear Radiation

I. Introduction

A. Nucleons

1. Neutrons and protons

B. Nuclides

1. Atoms identified by the number of protons and neutrons in the nucleus
 - a. radium-228 or ${}_{88}^{228}\text{Ra}$

II. Radioactivity

A. Radioactive Decay

1. The spontaneous disintegration of a nucleus into a slightly lighter and more stable nucleus, accompanied by emission of particles, electromagnetic radiation, or both

B. Nuclear Radiation

1. Particles or electromagnetic radiation emitted from the nucleus during radioactive decay

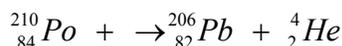
C. Unstable Nuclides

1. All nuclides beyond atomic # 83 are unstable and radioactive

III. Types of Radioactive Decay

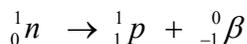
A. Alpha Emission

1. Alpha particle (α) is a helium nucleus (${}^4_2\text{He}$), so it has a 2+ charge.
2. Alpha emission is restricted almost entirely to very heavy nuclei

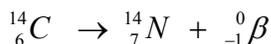


B. Beta Emission

1. Beta particle (β) is an electron emitted from the nucleus during nuclear decay

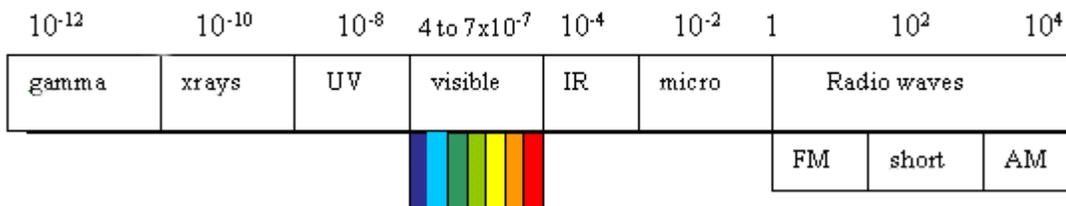


2. Beta particles are emitted when a neutron is converted into a proton and an electron



C. Gamma Emission

1. Gamma rays (γ) are high-energy electromagnetic waves emitted from a nucleus as it changes from an excited state to a ground energy state
2. Gamma rays are produced when nuclear particles undergo transitions in energy levels
3. Gamma emission usually follows other types of decay that leave the nucleus in an excited state

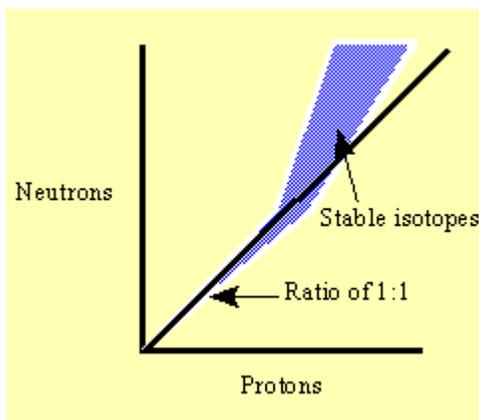


25.2 Radioactive Decay

I. Nuclear Stability and Decay

A. Neutron-to-Proton Ratio determines the type of decay that occurs

1. Band of Stability



II. Half-Life

A. Half-Life ($t_{1/2}$)

1. The time required for half the atoms of a radioactive nuclide to decay
 - a. More stable nuclides decay slowly
 - b. Less stable nuclides decay rapidly

| Nuclide | Half-life | Nuclide | Half-life |
|-----------------------|-------------------------|------------------------|--------------------------|
| ^3_1H | 12.32 years | $^{214}_{84}\text{Po}$ | 163.7 μ seconds |
| $^{14}_6\text{C}$ | 5715 years | $^{218}_{84}\text{Po}$ | 3.0 minutes |
| $^{32}_{15}\text{P}$ | 14.28 days | $^{218}_{85}\text{At}$ | 1.6 seconds |
| $^{40}_{19}\text{K}$ | 1.3×10^9 years | $^{238}_{92}\text{U}$ | 4.46×10^9 years |
| $^{60}_{27}\text{Co}$ | 10.47 minutes | $^{239}_{94}\text{Pu}$ | 2.41×10^4 years |

III. Transmutation Reactions

A. Transmutations

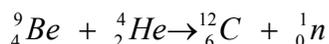
1. A change in the identity of a nucleus as a result of a change in the number of its protons

B. Nuclear Reaction

1. A reaction that affects the nucleus of an atom
2. Small amounts of mass are converted to LARGE amounts of energy
 - a. $E = mc^2$

C. Balancing Nuclear Reactions

1. Total atomic numbers and mass numbers must be equal on both sides



25.3 Fission and Fusion of Atomic Nuclei

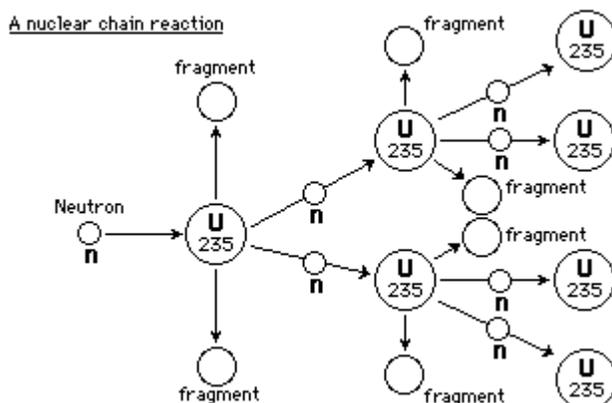
I. Nuclear Fission

A. Nuclear Fission

1. A very heavy nucleus splits into more stable nuclei of intermediate mass
2. The mass of the products is less than the mass of the reactants. Missing mass is converted to energy
 - a. Small amounts of missing mass are converted to HUGE amounts of energy ($E = mc^2$)

B. Nuclear Chain Reaction

1. A reaction in which the material that starts the reaction is also one of the products and can start another reaction



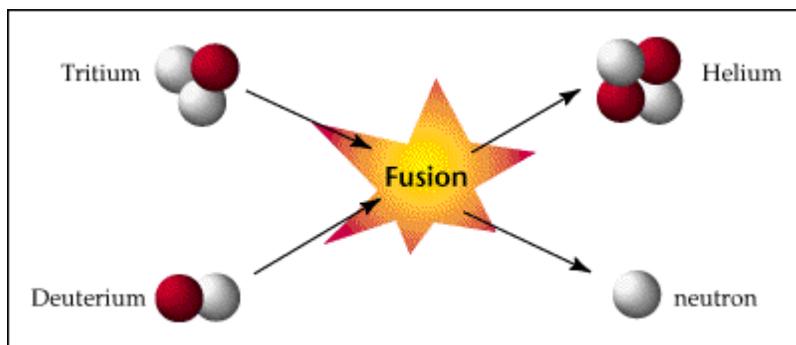
C. Critical Mass

1. The minimum amount of nuclide that provides the number of neutrons needed to sustain a chain reaction

II. Nuclear Fusion

A. Nuclear Fusion

1. Light-mass nuclei combine to form a heavier, more stable nucleus



B. Fusion Reactions

